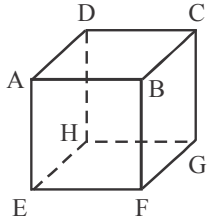
Part-A

1. Which of the following numbers is the largets?

$2^{3^4}, 2^{4^3}, 3^{2^4}, 3^{4^2}, 4^{2^3}, 4^{3^2}$

(a) 2^{3^4} (b) 3^{4^2} (c) 4^{3^2} (d) 4^{2^3}

2. The cube ABCDEFGH in the figure has each edge equal to a. The area of the triangle with vertices at A, C and F is:



(a) $\frac{\sqrt{3}}{4}a^2$ (b) $\frac{\sqrt{3}}{2}a^2$ (c) $\sqrt{3}\,a^2$ (d) $2\sqrt{3}\,a^2$

3. What is the number of distinct arrangements of the letters of the word UGCCSIR so that U and I cannot come together?

(a) 2520 (b) 720 (c) 1520 (d) 1800

4. Suppose the sum of the seven positive numbers is 21. What is the minimum possible value of the average of the squares of these numbers?

(a) 63 (b) 21 (c) 9 (d) 7

5. Let $A = \frac{1^{13} + 2^{13} + 3^{13} + ... + 100^{13}}{100}$, $B = \frac{1^{13} + 3^{13} + 5^{13} + ... + 99^{13}}{50}$, $C = \frac{2^{13} + 4^{13} + 6^{13} + ... + 100^{13}}{50}$

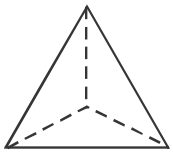
Which of the following is true?

(a) B < C < A (b) A < B < C (c) B < A < C (d) C < A < B

6. A circle of radius 5 units in the XY plane has its centre in the first quadrant, touches the x-axis and has a chord of length 6 units on the y-axis. The coordinates of its centre are

(a) 4, 6 (b) 3, 5 (c) 5, 4 (d) 4, 5

7. A wire of length 6cm is used to make a tetrahedron of each edge 1m, using only one strand of wire for each edge. The minimum number of times the wire has to be cut is

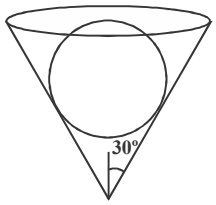


(a) 2 (b) 3 (c) 1 (d) 0

8. If the sum of the next two terms of the series below is x, what is the value of $\log_2 x$?

2, –4. 8, –16, 32, –64, 128, ……..

(a) 128 (b) 10 (c) 256 (d) 8

9. 

A conical vessel with semi-vertical angle 30º and height 10.5 cm has a thin lid. A sphere kept inside it touches the lid. The radius of the sphere in cm is

(a) 3.5 (b) 5 (c) 6.5 (d) 7

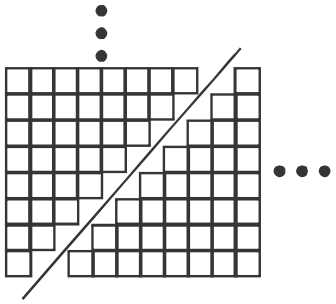
10. Amar, Akbar and Anthony are three friends, one of whom is a doctor, another is an engineer and the third is a professor. Amar is not an engineer. Akbar is the shortest. the tallest person is a doctor. The engineer's height is the geometric mean of the heights of the other two. Then which of the following is true?

(a) Amar is a doctor and he is the tallest
(b) Akbar is a professor and he is the tallest
(c) Anthony is an engineer and he is shortest
(d) Anthony is a doctor and he is the tallest

11. If 100 cats catch mice in 100 minutes, then how long will it take for 7 cats to catch 7 mice?

(a) 100/7 minutes (b) 100 minuts (c) 49/100 minutes (d) 7 minutes

12. What does this diagram demonstrate?



(a) $1+2+3+....+n=\frac{n(n+1)}{2}$ (b) $1^2+2^2+3^2+....+n^2=\frac{n(n+1)(2n+1)}{6}$

(c) $1+3+.....+(2n+=-1)=n^2$ (d) $2^2+4^2+....+(2n)^2=\frac{2.n(n+1)(2n+1)}{3}$

13. Suppose therea are socks of N different colors in box. If you take out one sock at a time, what is the maximum number of socks that you have to take out before a matching pair is found? Assume that N is na even number.

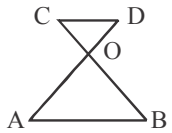
(a) N (b) N + 1 (c) N–1 (d) N/2

14. At what time after 4 O' clock, the hour and the minute hands will lie opposite to each other?

(a) $4-50'-30''$ (b) $4-52'-51''$ (c) $4-53'-23''$ (d) $4-54'-33''$

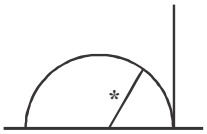
15. Which of the following curves just touches the 'x' axis?

(a) $y=x^2-x+1$ (b) $y=x^2-2x+2$ (c) $y=x^2-10x+25$ (d) $y=x^2-7x+12$

16. 

If AB is parallel to CD and AO = 2OD, then the area of triangle OAB is bigger than the area of triangle OCD by a factor of

(a) 1 (b) 3 (c) 4 (d) 8

17. 

A semi-circular arch of radius R has a vertical pole put on the ground together with one of its legs. An ant on the top of the arch finds the angular height of the tip of the pole to be 45°. The height of the pole is

(a) $\sqrt{2}$ R (b) $\sqrt{3}$ R (c) $\sqrt{4}$ R (d) $\sqrt{5}$ R

18. Suppose we make N identical smaller spheres from a big sphere. The total surface area of the smaller spheres is X times the total surface area of the big sphere, where X is

(a) $\sqrt{N}$ (b) 1 (c) $N^{1/3}$ (d) N^3.

19. What is the next number in the sequence 24, 30, 33., 39, 51, ...?
(a) 57 (b) 69 (c) 54 (d) 81

20. Four lines are drawn on a plane with no two parallel and no three concurrent. Lines are drawn joining the points of intersection of the previous four lines. The number of new lines obtained this way is:
(a) 3 (b) 5 (c) 12 (d) 2

Part-B

21. For an odd nucleon in 'g' nuclear orbital and parallel to I, spin and parity are
(a) 9/2 and (+) (b) 7/2 and (+) (c) 9/2 and (–) (d) 7/2 and (–)

22. For the deposition of Pb by electroplating, the best suited compound among the following is
(a) $PbCl_2$ (b) $PbSO_4$ (c) $Pb(Et)_4$ (d) $Pb(BF_4)_2$.

23. Appropriate reasons for the deviation form the Beer's law among the following are
(A) Monochromaticity of light (C) Very high concentration of analyte
(B) Association of analyte (D) Dissociation of analyte.
(a) A, B and D (b) B, C and D (c) A, C and D (d) A, B and C

24. Which one of the following shows the highest solubility in hot concentrated aqueous NaOH?
(a) $La(OH)_3$ (b) $Nd(OH)_3$ (c) $Sm(OH)_3$ (d) $Lu(OH)_3$.

25. In the vibrational spectrum of CO_2, the number of fundamental vibrational modes common in both infrared and Raman are
(a) Three (b) Two (c) One (d) Zero

26. The light pink color of $\left[Co\left(H_2O\right)_6\right]^{2+}$ and the deep blue color of $\left[CoCl_4\right]^{2-}$ are due to
(a) MLCT transition in the first and d-d transition in the second
(b) LMCT transition in both
(c) d-d transitions in both
(d) d-d transition in the first and MLCT transition in the second.

27. In $\left[Mo_2\left(S_2\right)_6\right]^{2-}$ cluster the number of birdging S_2^{2-} and coordination number of Mo respectively, are
(a) 2 and 8 (b) 2 and 6 (c) 1 and 8 (d) 1 and 6

28. ^1H NMR spectrum of HD would show
(a) a singlet (b) a doublet
(c) a triplet with intensity ratio 1:2:1 (d) a triplet with intensity ratio 1:1:1

29. The number of possible isomers of $\left[Ru\left(PPh_3\right)_2\left(acac\right)_2\right]$ (acac = acetylacetonate) is:
(a) 2 (b) 3 (c) 4 (d) 5

30. The total number of Cu–O bonds present in the crystalline copper(II) acetate monohydrate is:
(a) 10 (b) 6 (c) 8 (d) 4

31. The electronegativity differences is the highest for the pair
(a) Li, Cl (b) K, F (c) Na, Cl (d) Li, F

32. Which ones among CO_3^{2-}, SO_3, XeO_3 and NO_3^- and NO_3^- have planar structure?

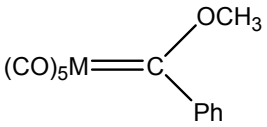
(a) CO_3^{2-}, SO_3 and XeO_3 (b) SO_3, XeO_3 and NO_3^-

(c) CO_3^{2-}, XeO_3 and NO_3^- (d) CO_3^{2-}, SO_3 and NO_3^-

33. The substitution of $\eta^5 - Cp$ group with nitric oxide is the easiest for

(a) $\eta^5 - Cp_2Fe$ (b) $\eta^5 - Cp_2CoCl$ (c) $\eta^5 - Cp_2Ni$ (d) $\eta^5 - Cp_2Co$

34. The molecule



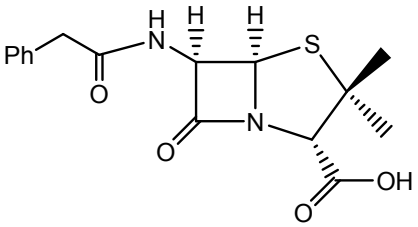
obeys 18 e rule. The two 'M' satisfying the condition are

(a) Cr, Re^+ (b) Mo, V (c) V, Re^+ (d) Cr, V

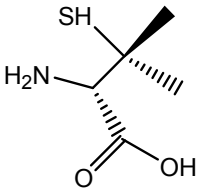
35. The correct ser of the biologically essential elements is,
(a) Fe, Mo, Cu, Zn (b) Fe, Cu, Co, Ru (c) Cu, Mn, Zn, Ag (d) Fe, Ru, Zn, Mg

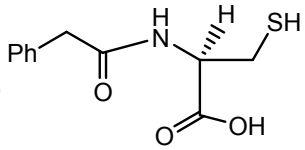
36. The number of lines exhibited by a high resolution EPR spectrum of the species, [Cu(ethylenediamine)2]2+ is [Nuclear spin (I) of Cu = 3/2 and that of N = 1]
(a) 12 (b) 15 (c) 20 (d) 36

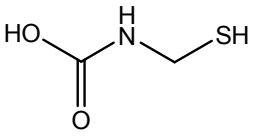
37. Degradation of penicillin G

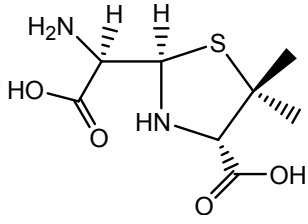


gives penicillamine that can utilize nitrogen, oxygen or sulfur atoms as donors to bind with lead(II), mercury (II) or copper(II). The structure of penicillamine is

(a)


(b)


(c)


(d)


38. The molecular that has an S_6 symmetry element is
(a) B_2H_6 (b) CH_4 (c) PH_5 (d) SF_6

39. The electric dipole allowed transition in a d^2 atomic system is
(a) $^3F \rightarrow {}^1D$ (b) $^3F \rightarrow {}^1P$ (c) $^3F \rightarrow {}^3D$ (d) $^3F \rightarrow {}^3P$

40. When a hydrogen atom is placed in an electric field along the y-axis, the orbital that mixes most with the ground state 1s obrital is
(a) 2s (b) $2p_x$ (c) $2p_y$ (d) $2p_x$

41. For water, $\Delta H_{vap} \approx 41$ kJ mol^{-1}. The molar entropy of vaporization at 1 atm pressure is approximately
(a) 410 J K^{-1} mol^{-1} (b) 110 J K^{-1} mol^{-1} (c) 41 J K^{-1} mol^{-1} (d) 11 JK^{-1} mol^{-1}.

42. If A and B are non-commuting hermitian operators, all eigenvalues of the operator given by the commutator [A, B] are
(a) complex (b) real (c) imaginary (d) zero

43. The value of commutator $\left[x, p_x^2\right]$ is given by
(a) 2i (b) $2i\hbar$ (c) $2i\hbar x$ (d) $2i\hbar\, p_x$

44. The correlation coefficient between two arbitrary variables x and y is zero, if
(a) $\langle xy\rangle = \langle yx\rangle$ (b) $\langle x^2\rangle = \langle x\rangle^2$ (c) $\langle y^2\rangle = \langle y\rangle^2$ (d) $\langle xy\rangle = \langle x\rangle\langle y\rangle$

45. A carnot takes up 90 J of heat from the source kept at 300K. The correct statement among the following is
(a) It transfers 60 J of heat to the sink at 200K
(b) It transfers 50 J of heat to the sink at 200K
(c) It transfers 60 J of heat to the sink at 250 K
(d) It transfers 50 J of heat to the sink at 250 K

46. The relative population in two states with energies E_1 and E_2 satisfying Boltzmann distribution is given by $n_1/n_2 = (3/2)\exp\left[-(E_1 - E_2)/k_B T\right]$. The relative degeneracy g_2/g_1 is:
(a) 2 (b) 2/3 (c) 3/2 (d) 3

47. The Daniel cell is
(a) $Pt_I(s)\,|\,Zn(s)\,|\,Zn^{2+}(aq)\,||\,Cu^{2+}(aq)\,|\,Cu(s)\,|\,Pt_{II}(s)$
(b) $Pt_I(s)\,|\,Zn(s)\,|\,Zn^{2+}(aq)\,||\,Ag^{+}(aq)\,|\,Ag(s)\,|\,Pt_{II}(s)$
(c) $Pt_I(s)\,|\,Fe(s)\,|\,Fe^{2+}(aq)\,||\,Cu^{2+}(aq)\,|\,Cu(s)\,|\,Pt_{II}(s)$
(d) $Pt_I(s)\,|\,H_2(s)\,|\,H_2SO_4(aq)\,||\,Cu^{2+}(aq)\,|\,Cu(s)\,|\,Pt_{II}(s)$

48. If the concept of half-life is generalized to quarter-life of a first order chemical reaction, it will be equal to
(a) $\ell n\ 2/k$ (b) $\ell n\ 4/k$ (c) $4/k$ (d) $1/4k$

49. Kohlrausch's law is applicable to a dilute solution of
(a) Potassium chloride in hexane (b) Acetic acid in water
(c) Hydrochloric acid in water (d) Benzoic acid in benzene

50. A dilute silver nitrate solution is added to a slight excess iodide solution. A solution of AgI is formed whose surface adsorbs.
(a) I^- (b) NO_3^- (c) Na^+ (d) Ag^+

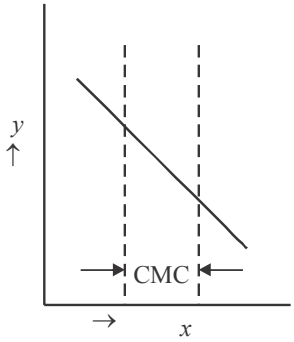
51. The absorption spectrum of O_2 shows a vibrational structure that becomes continuum at 56875 cm^{-1}. At the continuum, it dissociates into one ground state atom (O_g) and one excited state atom (O_e). The energy difference between O_e and O_g is 15125 cm^{-1}. The dissociation energy (in cm^{-1}) of ground state of O_2 is:

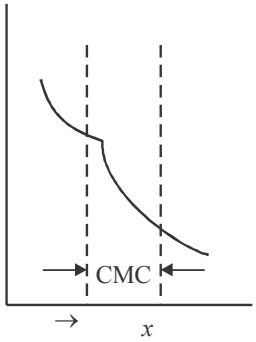
(a) $\frac{56875}{15125}$ (b) $\frac{15125}{56875}$ (c) 72000 (d) 41750

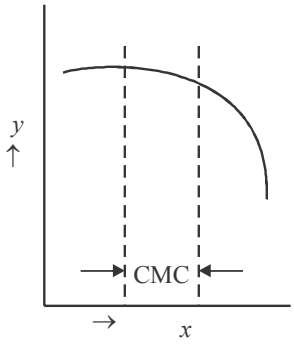
52. The angle between the two planes represented by the Miller indices (1 1 0) and (1 1 1) in a simple cubic lattice is:

(a) 30° (b) 45° (c) 60° (d) 90°

53. The correct representation of the variation of molar conductivity (y-axis) with surfactant concentration (x-axis) is [CMC = critical micelle concentration].

(a)


(b)


(c)


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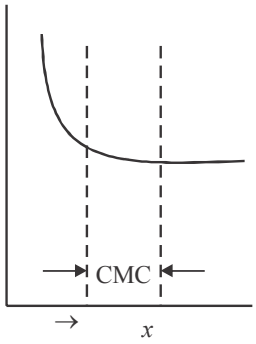
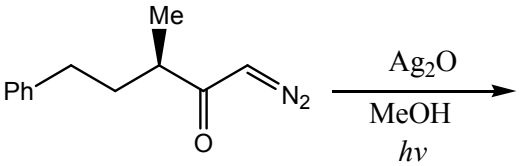
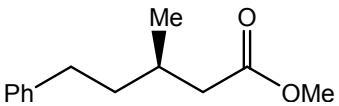
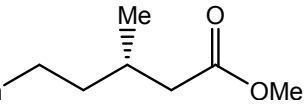
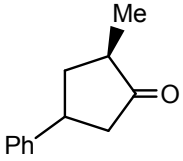


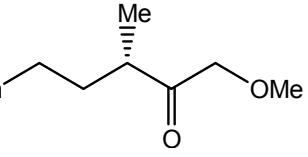
54. The major product formed in the following reaction is



(a)


(b)


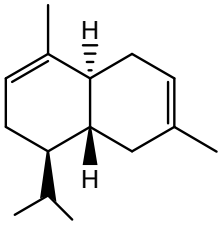
(c)


(d)


55. If the pK_a value for p-methoxybenzoic acid is 4.46 and that of benzoic acid is 4.19, the σ_{para} for methoxy group is:

(a) 8.65 (b) 4.32 (c) 0.27 (d) – 0.27

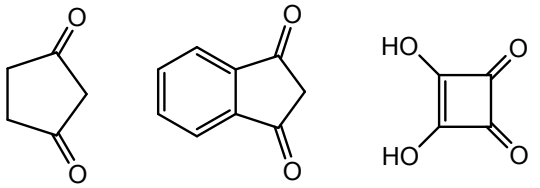
56. The biosynthetic precursor of cadinene is:



cadinene

(a) shikimic acid (b) mevalonic acid (c) arachidonic acid (d) prephenic acid.

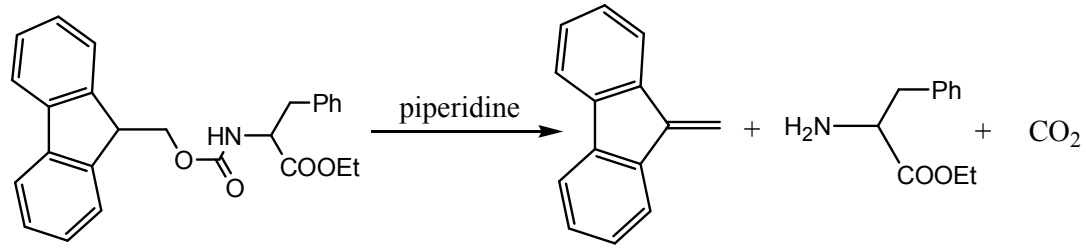
57. The correct order of acidity of the compounds A – C is:



A **B** **C**

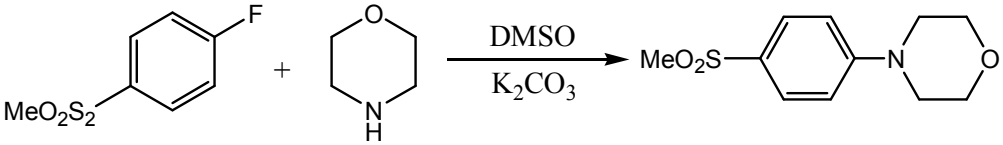
(a) A > B > C (b) B > C > A (c) C > A > B (d) B > A > C

58. The mechanism involved in the following conversion is:



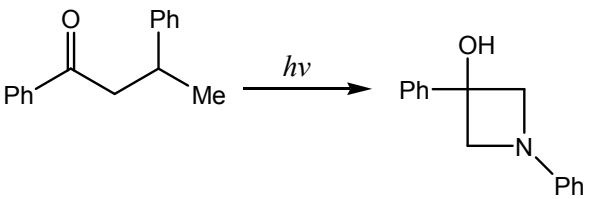
(a) E_2-elimination (b) E_1-elimination (c) syn-elimination (d) E_1 cb-elimination.

59. The correct statement(s)-A-D are given for the following reaction. The correct one(s) is (are)



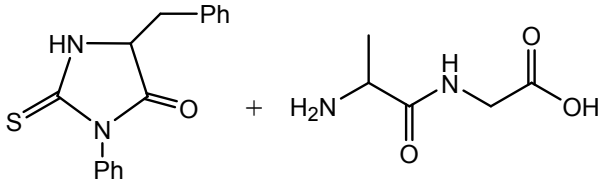
(a) aromatic ipso substitution reaction (b) aromatic nucleophilic substitution
(c) aromatic electrophilic substitution (d) aromatic free radical substitution.

60. The following photochemical transformation proceeds through



(a) Norrish type I reaction (b) Norrish type II reaction
(c) Barton reaction (d) Paterno-Buchi reaction

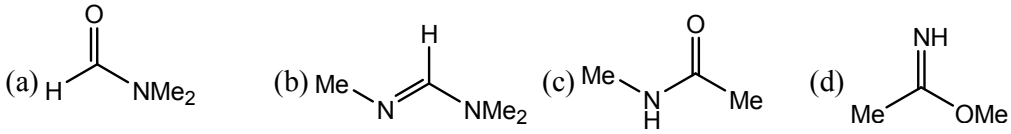
61. A tripeptide gives the following products on Edman degradation.



The tripeptide is

(a) Phe-Ala-Gly (b) Phe-Gly-Ala (c) Ala-Gly-Phe (d) Gly-Ala-Phe

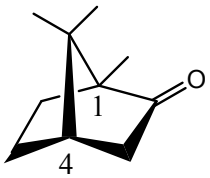
62. In the ^{1}H NMR spectrum recorded at 293 K, an organic compound (C_3H_7NO), exhibited signals at δ 7.8 (1H, s), 2.8 (3H, s) and 2.6 (3H, s). The compound is



63. In the IR spectrum of p-nitrophenyl acetate, the carbonyl absorption band appears at

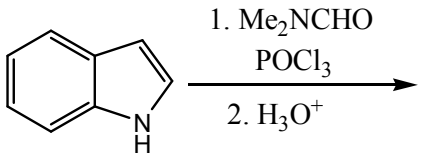
(a) 1670 cm^{-1} (b) 1700 cm^{-1} (c) 1730 cm^{-1} (d) 1760 cm^{-1}.

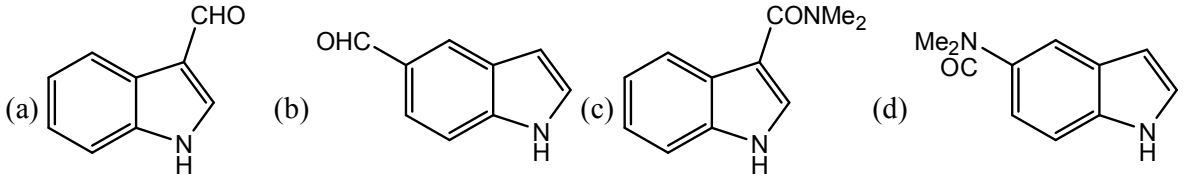
64. The absolute configuration at the two chiral centres of (–)–camphore is:



(a) 1R, 4R (b) 1R, 4S (c) 1S, 4R (d) 1S, 4S

65. The major product formed in the following reaction is





66. The first person to separate a racemic mixture into individual enantiomers is

(a) J, H van't Hoff (b) Pasteur (c) H.E. Fischer (d) F. Wohler

67. Consider the following statements for [18]-annulene

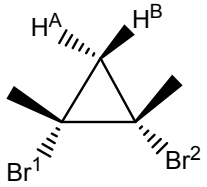
(A) It is aromatic

(B) The inner protons resonate at δ 9.28 in its ^{1}H NMR spectrum

(C) There are six protons in the shielded zone.

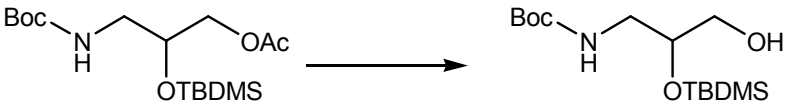
(a) A, B, C (b) A and B only (c) B and C only (d) A and C only

68. In the compound give below, the relation between H^A, H^B; and between Br^1, Br^2 is:



(a) H^A, H^B are enantiotropic; and Br^1 Br^2 are diastereotopic
(b) H^A, H^B are diastereotopic; and Br^1, Br^2 are enantiotropic
(c) H^A, H^B are diastereotopic; and Br^1, Br^2 are homotopic
(d) H^A, H^B are enantiotropic; and Br^1, Br^2 are homotopic.

69. The most appropriate reagent to effect the following chemoselective conversion is



(a) HCl, EtOH, reflux (b) Bu_4NF
(c) K_2CO_3, MeOH (d) CF_3COOH, EtOH, rt.

70. Among the following, an example of a "Green Synthesis" is
(a) Synthesis of malachite green
(b) Friedel-Craft's acylation of anisole with Ac_2O/anhydrous $AlCl_3$.
(c) Jones' oxidation of benzyl alcohol to benzoic acid.
(d) Diels-Alder reaction of furan and maleic acid in water.

Part-C

71. The recoil energy of a Mossabauer nuclide of mass 139 amu is 2.5 MeV. The energy emitted by the nucleus in keV is:
(a) 12.5 (b) 15.0 (c) 20.5 (d) 25.0

72. Complexes of general formula, fac-$[Mo(CO)_3(phosphite)_3]$ have the C—O stretching bands as given below.
Phosphines: PF_3(A); PCl_3(B); $P(Cl)Ph_2$(C); PMe_3(D)
ν(CO), cm^{-1}: 2090(i); 2040(ii); 1977(iii); 1945(iv)
The correct comibination of the phsphine and the streching frequency is,
(a) (A–i), (B–ii), (C–iii), (D–iv) (b) (A–ii), (B–i), (C–iv), (D–iii)
(c) (A–iv), (B–iii), (C–ii), (D–i) (d) (A–iii), (B–iv), (C–i), (D–ii)

73. On subjecting 9.5 ml solution of Pb^{2+} of X M to polorographic measurements, I_d was found to be 1 μA. When 0.5 mL of 0.04 M Pb^{2+} was added before the measurement, the Id was found to be 1.25 μA.
(a) 0.0035 (b) 0.0400 (c) 0.0067 (d) 0.0080

74. Match each item from the **List-I** (compound in solvent) with that from the **List-II** (its behaviour) and select the correct combination using the codes given below.

List-I	**List-II**
A. CH_3COOH in pyridine	(i) strong acid.
B. CH_3COOH in H_2SO_4	(ii) weak acid
C. $HClO_4$ in H_2SO_4	(iii) strong base
D. SbF_5 in HF	(iv) weak base

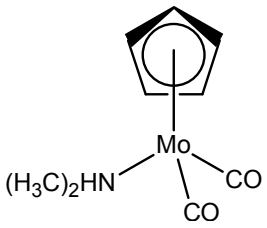
(a) (A–i), (B–ii), (C–iii), (D–iv) (b) (A–ii), (B–i), (C–iii), (D–iv)
(c) (A–iii), (B–iv), (C–ii), (D–i) (d) (A–iv), (B–ii), (C–iii), (D–i)

75. Structure of a carborane with formula, $C_2B_4H_8$ is formally derived from
(a) Closo-borane (b) Nido-borane (c) Arachno-borane (d) Conjuncto-borane

76. Boric acid is a weak acid in aqueous solution. But its acidity increases significantly in the presence of ethylene glycol, because
(a) ethylene glycol releases additional H^+
(b) $B(OH)_4^-$ is consumed in forming a compound with ethylene glycol.
(c) ethylene glycol neutralizes H^+ released by boric acid.
(d) Boric acid dissociates better in the mixed-solvent.

77. Coordination number of "C" in Be_2C_3 whose structure is correlated with that of CaF_2, is:
(a) 2 (b) 4 (c) 6 (d) 8

78. For the molecule below,



consider the following statements about its room temperature spectral data.
(A) 1H NMR has singlets at 5.48 and 3.18 ppm
(B) 1H NMR has multiplet at 5.48 and singlet at 3.18 ppm
(C) IR has CO stretching bands at 1950 and 1860 cm^{-1}
(D) IR has only one CO stretching band at 1900 cm^{-1}.
The correct pair of statement is,
(a) A and C (b) B and C (c) A and D (d) B and D

79. In the cluster $\left[Co_3(CH)(CO)_9\right]$ obeying 18e rule, the number of metal-metal bonds and the bridgind ligands respectively, are
(a) 3 and 1 CH (b) 0 and 3 CO (c) 3 and 1 CO (d) 6 and 1 CH

80. Consider the ions Eu(III), Gd(III), Sm(II) and Lu(III). The observed and calculated magnetic moment values are closest for the pair
(a) Gd(III), Lu(III) (b) Eu(III), Lu(III) (c) Sm(III), Gd(III) (d) Sm(III), Eu(III)

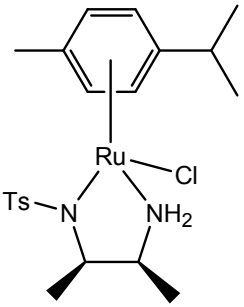
81. Silicates with continuous 3D frame work are
(a) Neso-silicates (b) Soro-silicates (c) Phyllo-silicates (d) Tecto-silicates

82. The correct spinel structure of Co_3O_4 is:
(a) $\left(Co^{2+}\right)_t\left(2Co^{3+}\right)_0 O_4$ (b) $\left(Co^{2+}\right)_t\left(2Co^{3+}Co^{3+}\right)_0 O_4$
(c) $\left(Co^{2+}Co^{3+}\right)_t\left(Co^{3+}\right)_0 O_4$ (d) $\left(2Co^{3+}\right)_t\left(Co^{2+}\right)_0 O_4$

83. In the solid state, the $CuCl_5^{3-}$ ion has two types of bonds. These are
(a) Three long and two short (b) Two long and three short
(c) One long and four short (d) Four long and one short

84. In metalloenzymes, the metal centres are covalently linked through the side chains of the amino acid residues. The correct set of amino acids which are involved in the primary coordinates spheres of metalloenzymes is
(a) Ala, Leu, His (b) Glu, His, Cys (c) Leu, Glu, Cys (d) Ala, His, Glu

85. Consider the catalyst in **column-I** and reactin in **column-II**

Column-I	**Column-II**
A. [(R)-BINAP]Ru^{2-}	(i) hydroformylation
B. $[Rh(CO)_2I_2]^-$	(ii) asymmetric hydrogenation.
C. $Pd(PPh_3)_4$	(iii) asymmetric hydrogen transfer
D. 	(iv) heck coupling.

The best match of a catalyst of column-I with the reaction nuclear column-II is

(a) (A–ii), (B–i), (C–iv), (D–iii) (b) (A–i), (B–ii), (C–iii), (D–iv)

(c) (A–iii), (B–i), (C–iv), (D–ii) (d) (A–iv), (B–iii), (C–ii), (D–i)

86. A solution of 2.0 g of brass was analysed for Cu electrogravimetrically using Pt-gauze as electrode. The weight of Pt-gauze changed from 14.5g to 16.0 g. The percentage weight of Cu in brass is

(a) 50 (b) 55 (c) 60 (d) 75

87. The platinum complex of NH_3 and Cl^- ligands is an anti-tumour agent. The correct isomeric formula of the complex and its precursor are

(a) cis-$Pt(NH_3)_2Cl_2$ and $PtCl_4^{2-}$ (b) trans-$Pt(NH_3)_2Cl_2$ and $PtCl_4^{2-}$

(c) cis-$Pt(NH_3)_2Cl_2$ and $Pt(NH_3)_4^{2+}$ (d) trans-$Pt(NH_3)_2Cl_2$ and $Pt(NH_3)_4^{2-}$

88. Successive addition of NaCl, H_3PO_4, KSCN and NaF to a solutin of $Fe(NO_3)_3.9H_2O$ gives yellow, colourless, red and again colorless solutions due to the respective formation of:

(a) $\left[Fe(H_2O)_5Cl\right]^{2+}, \left[Fe(H_2O)_5(PO_4)\right], \left[Fe(H_2O)_5(SCN)\right]^{2+}, \left[Fe(H_2O)_5F\right]^{2+}$

(b) $\left[Fe(H_2O)_4Cl(OH)\right]^{+}, \left[Fe(H_2O)_5(PO_4)\right], \left[Fe(H_2O)_5(SCN)\right]^{2-}, \left[Fe(H_2O)_5F\right]^{2+}$

(c) $\left[Fe(H_2O)_5(Cl)\right]^{2+}, \left[Fe(H_2O)_6\right]^{3+}, \left[Fe(H_2O)_5(SCN)\right]^{2+}, \left[Fe(H_2O)_5F\right]^{2+}$

(d) $\left[Fe(H_2O)_5Cl\right]^{2+}, \left[Fe(H_2O)_5(PO_4)\right], \left[Fe(H_2O)_5(SCN)\right]^{2+}, \left[Fe(H_2O)_4(SCN)F\right]^{+}$

89. Which one of the following will NOT undergo oxidative addition by methyl iodide?

(a) $\left[Rh(CO)_2I_2\right]^-$ (b) $\left[Ir(PPh_3)_2(CO)Cl\right]$

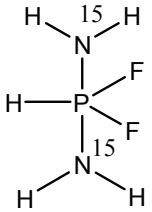
(c) $\left[\eta^2-CpRh(CO)_2\right]$ (d) $\left[\eta^5-Cp_2Ti(Me)Cl\right]$

90. In hydrofomylation reaction using $\left[Rh(PPh_3)_3(CO)(H)\right]$ as the catalyst, addition of excess PPh_3 would

(a) increase the rate of reaction (b) decrease the rate of reaction.

(c) not influence of the rate of reaction (d) stop the reaction.

91. Find out the number of lines in the ^{31}P NMR signal for



(a) 3 (b) 6 (c) 18 (d) 90

92. The rate of exchange of OH_2 present in the coordination sphere by $^{18}OH_2$ of, (i) $[Cu(OH)_2)_6]^{2+}$, (ii) $[Mn(OH_2)_6]^{2+}$, (iii) $Fe(OH_2)_6]^{2+}$, (iv) $[Ni(OH_2)_6]^{2+}$, follows an order

(a) (i) > (ii) > (iii) > (iv) (b) (i) > (iv) > (iii) > (ii)

(c) (ii) > (iii) > (iv) > (i) (d) (iii) > (i) > (iv) > (ii)

93. Based on the behaviour of the metalloenzymes, consider the following statements

(A) In the enzymes, the zinc activates O_2 to form peroxide species.

(B) In the enzymes, the zinc activates H_2O and provides a zinc boud hydroxide.

(C) In the oxidases, the iron activates O_2 to break the bonding between the two oxygens

(D) Zinc ion acts as a nucleophile and attacks at the peptide carbonyl

The set of correct statements is,

(a) A and B (b) B and C (c) C and D (d) A and D

94. Fe^{2+}-porphyrins fail to exhibit reversible oxygen transport and cannot differentiate CO from O_2. However, the hemoglobin is free from both these pit falls. Among the following

(A) Fe^{2+} - porphyrins undergo μ-oxodimer formation and the same is prevented in case of the hemoglobin.

(B) Fe–CO bond strength is much low in case of hemoglobin when compared to the Fe^{2+} - porphyrins.

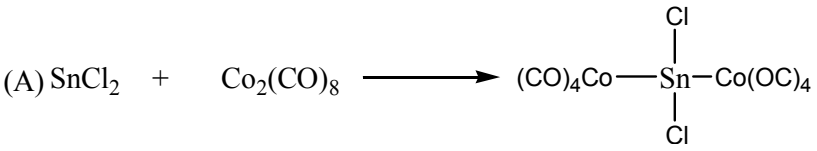
(C) While Fe–CO is linear, Fe–O_2 is bent and is recognized by hemoglobin

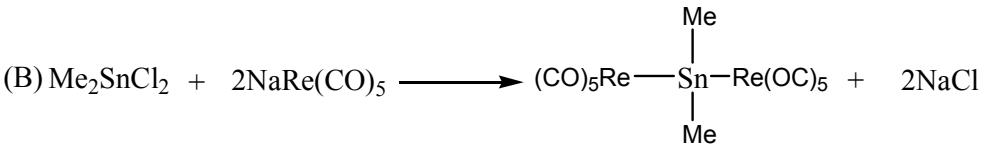
(D) The interlinked four monomeric units in the hemoglobin are responsible to overcome the pitfalls.

The correct set of statements is

(a) A and B (b) A and C (c) C and D (d) B and D

95. Reactions A and B are, termed as respectively.





(a) Insertion, Metathesis (b) Metathesis, insertion

(c) Oxidative, addition, metathesis (d) Oxidative addition, insertion

96. A metal crystallizes in fcc structure with a unit cell side of 500 pm. If the density of the crystal is 1.33 g/cc, the molar mass of the metal is close to

(a) 23 (b) 24 (c) 25 (d) 26

97. The activation energy for the bimolecular reaction $A + BC \rightarrow AB + C$ is E_0 in the gas phase. If the reaction is carried out in a confined volume of λ^3, the activation energy is expected to

(a) remain unchanged (b) increase with decreasing λ.

(c) decrease with decreasing λ. (d) oscillate with decreasing λ.

98. In a many-electron atom, the total orbital angular momentum (L) and spin (S) are good quantum numbers instead of the individual orbital (l_1, l_2) and spin (s1, s2) angular momenta in the presence of
(a) inter-electron repulsion (b) spin-orbit interaction
(c) hyperfine coupling (d) external magnetic field.

99. The packing fraction of a simple cubic lattice is close to
(a) 0.94 (b) 0.76 (c) 0.52 (d) 0.45

100. The number of IR active vibrational modes of pyridine is:

C_{2v}	E_2	C_2	σ_v	σ_v'	
A_1	1	1	1	1	z
A_2	1	1	−1	−1	R_z
B_1	1	−1	1	−1	x, R_y
B_2	1	−1	−1	1	y, R_x

(a) 12 (b) 20 (c) 24 (d) 33

101. One of the excited states of Ti has the electric configuration $[Ar]\,4s^2 3d^1 4p^1$. The number of microstates with zero total spin (S) for this configuration is
(a) 9 (b) 15 (c) 27 (d) 60

102. For the reaction $A_2 \rightleftharpoons 2A$ in a closed container, the relation between the degree of dissociation (α) and the equilibrium constant K_p at a fixed temperature is given by

(a) $\alpha = \left[K_p/\left(K_p + 4p\right)\right]$ (b) $\alpha = \left[K_p/\left(K_p + 4p\right)\right]^{1/2}$

(c) $\alpha = \left[\left(K_p + 4p\right)/K_p\right]$ (d) $\alpha = \left[\left(K_p + 4p\right)/K_p\right]^{1/2}$

103. The fugacity of a gas depends on pressure and the compressibility factor $Z\left(= p\bar{V}/RT\right)$ through the relation $\left(\bar{V} \text{ is the molar volume}\right)$
For most gases at temperature T and up to moderate pressure, this equation shows that
(a) $f < p$, if $T \to 0$ (b) $f < p$, if $T \to \infty$
(c) $f > p$, if $T \to 0$ (d) $f = p$, if $T \to 0$

104. The internal pressure $\left(\partial U/\partial V\right)_T$ of a real gas is related to the compressibility factor $Z = p\bar{V}/RT$ by $\left[\bar{V} \text{ is the molar volume}\right]$

(a) $\left(\partial U/\partial V\right)_T = RT\left(\partial Z/\partial V\right)_T$ (b) $\left(\partial U/\partial V\right)_T = RT/\left(\bar{V}\,Z\right)$

(c) $\left(\partial U/\partial V\right)_T = \left(RT^2/\bar{V}\right)\left(\partial Z/\partial V\right)_V$ (d) $\left(\partial U/\partial V\right)_T = \left(\bar{V}/RT^2\right)\left(\partial Z/\partial T\right)_V$

105. Suppose, the ground stationary state of a harmonic oscillator with force constant 'k' is given by

$$\psi_0 = \exp\left[-Ax^2\right]$$

Then, A should depend on k as
(a) $A \propto k^{-1/2}$ (b) $A \propto k$ (c) $A \propto k^{1/2}$ (d) $A \propto k^{1/3}$

106. Combining two real wave functions ϕ_1 and ϕ_2, the following functions are constructed: $A = \phi_1 + \phi_2$, $B = \phi_1 + i\phi_2$, $C = \phi_1 - i\phi_2$, $D = i(\phi_1 + \phi_2)$. The correct statement will then be

(a) A and B represent the same state (b) A and C represent the same state.
(c) A and D represents the same state (d) B and D represent the same state.

107. Crystal A diffracts from (1 1 1) and (2 0 0) planes but not from (1 1 0) plane, while the crystal B diffracts from (1 1 0) and (2 0 0) planes but not from the (1 1 1) plane. From the above, we may conclude that

(a) A has fcc lattice while B has bcc lattice (b) A has bcc lattice while B has fcc lattice
(c) A and B both have fcc lattice (d) A and B both have bcc lattice.

108. The decomposition of NH_3 on Mo surface follows Langmuir-Hinshelwood mechanism. The decomposition was carried out at low pressures. The initial pressure of NH_3 was 10^{-2} torr. The pressure of NH_3 was reduced to 10^{-4} torr in 10 minutes. The rate constant of decomposition of NH_3 is:

(a) 9.9×10^{-4} torr min^{-1} (b) 0.4606 min^{-1}

(c) 9.9×10^{-3} torr min^{-1} (d) 0.693 min^{-1}

109. A polymer sample has the following composition.

Number of molecules	Molecular weight
10	1000
50	2000
40	4000

The polydispersity index (P.D.I) of the polymer is

(a) $\frac{85000}{27}$ (b) $\frac{85}{81}$ (c) $\frac{850}{729}$ (d) $\frac{729}{850}$

110. The equilibrium constant for an electrochemical reaction,

$$2Fe^{3+} + Sn^{2+} \rightleftharpoons 2Fe^{2+} + Sn^{4+}$$

is $\left[E^0\left(Fe^{3+}/Fe^{2+}\right) = 0.75\ V,\ E^0\left(Sn^{4+}/Sn^{2+}\right) = 0.15V,\ (2.303RT/F) = 0.06V\right]$

(a) 10^{10} (b) 10^{20} (c) 10^{30} (d) 10^{40}.

111. A bacterial colony grows most commonly by cell division. The change in the population due to cell division in an actively growing colony is $dN = \lambda_g N\, dt$. The population of bacterial colony at time 't' is $\left[N_0 = N(t=0)\right]$

(a) $N_0\lambda_g t$ (b) $N_0 \exp\left[-\lambda_g t\right]$ (c) $N_0 \exp\left[\lambda_g t\right]$ (d) $N_0(\lambda_g t)^2$

112. The Arrhenius parameters for the thermal decomposition of NOCl, $2NOCl(g) \rightarrow 2NO(g) + Cl_2(g)$, are $A = 10^{13} M^{-1}s^{-1}$, $E_a = 105$ kJ mol^{-1} and $RT = 2.5$ kJ mol^{-1}. The enthalpy (in kJ mol^{-1}) of the activated complex will be

(a) 110 (b) 105 (c) 102.5 (d) 100

113. The rotational partition function of H_2 is:

(a) $\sum_{J=0,\,1,\,2.....} (2J+1)\, e^{-\beta hcBJ(J+1)}$

(b) $\sum_{J=1,3,\,5,.....} (2J+1)\, e^{-\beta hcBJ(J+1)}$

(c) $\sum_{J=0,2,4.....} (2J+1)\, e^{-\beta hcBJ(J+1)}$

(d) $\frac{1}{4}\sum_{J=0,2,4.....} (2J+1)\, e^{-\beta hcBJ(J+1)} + 3\sum_{J=1,3,5.....} (2J+1)\, e^{-\beta hcBJ(J+1)}$

114. The potential in Debye-Huckel theory is proportional to

(a) $1/\kappa r$ (b) $\exp[-\kappa r]$ (c) $\exp[-\kappa r]/r$ (d) κr

115. The vibrational frequency and anharmonicity constant of an alkali halide are 300 cm^{-1} and 0.0025 respectively. The positions (in cm^{-1}) of its fundamental mode and first overtone are respectively.

(a) 300, 600 (b) 298.5, 595.5 (c) 301.5, 604.5 (d) 290, 580

116. The adsorption of a gas is described by the Langmuir isotherm with the equilibrium constant $K = 0.9\ kPa^{-1}$ at 25ºC. The pressure (in kPa) at which the fractional surface coverage is 0.95, is

(a) 1/11.1 (b) 21.1 (c) 11.1 (d) 42.2

117. The energy of a harmonic oscillator in its ground state is $\frac{1}{2}\hbar\omega$. According to the virial theorem, the average kinetic (T) and potential (V) energies of the above are

(a) $T = \frac{1}{4}\hbar\omega;\quad V = \frac{1}{4}\hbar\omega$ (b) $T = \frac{1}{8}\hbar\omega;\quad V = \frac{3}{8}\hbar\omega$

(c) $T = \hbar\omega;\quad V = -\frac{1}{2}\hbar\omega$ (d) $T = \frac{3}{8}\hbar\omega;\quad V = \frac{1}{8}\hbar\omega$

118. The energy of a hydrogen atom in a state is $-\frac{hcR_H}{25}\left(R_H = \text{Rydberg constant}\right)$. The degeneracy of the state will be

(a) 5 (b) 10 (c) 25 (d) 50

119. The trial wave function of a system is expanded as $\psi_t = c_1\varphi_1 + c_2\varphi_2$. The matrix elements of the Hamiltonian are $\langle\varphi_1|H|\varphi_1\rangle = 0;\ \langle\varphi_1|H|\varphi_2\rangle = 2.0 = \langle\varphi_2|H|\varphi_1\rangle$ and $\langle\varphi_2|H|\varphi_2\rangle = 3.0$. The approximate ground-state energy of the system from the linear variational principle is

(a) –1.0 (b) –2.0 (c) +4.0 (d) +5.0

120. One molecular orbital of a polar molecule AB has the form $c_A\psi_A + c_B\psi_B$, where ψ_A and ψ_B are normalized atomic oribitals centred on A and B, respectively. The electron in this orbital is found on atom B with a probability of 90%. Neglecting the overalp between ψ_A and ψ_B, a possible set of c_A and c_B is:

(a) $c_A = 0.95, c_B = 0.32$ (b) $c_A = 0.10, c_B = 0.90$

(c) $c_A = -0.95, c_B = 0.32$ (d) $c_A = 0.32, c_B = 0.95$

121. 4-Hydroxybenzoic acid exhibited signals at δ 171, 162, 133, 122 and 116 ppm in its broadband decoupoled ^{13}C NMR spectrum. The correct assignment of the signals is

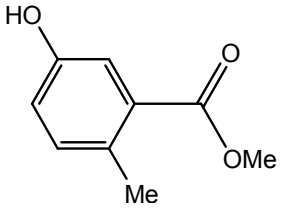
(a) $\delta\ 171(C-4), 162(COOH), 133(C-3\ \&\ 5), 122(C-1)$ and $116(C-2\ \&\ 6)$

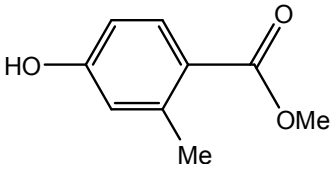
(b) $\delta\ 171(COOH), 162(C-4), 133(C-2\ \&\ 6), 122(C-1)$ and $116(C-3\ \&\ 5)$

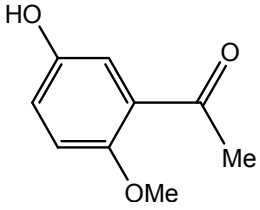
(c) $\delta\ 171(C-4), 162(COOH), 133(C-2\ \&\ 6), 122(C-1)$ and $116(C-3\ \&\ 5)$

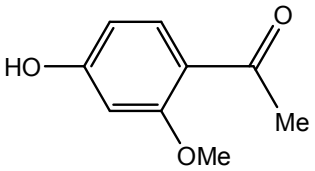
(d) $\delta\ 171(COOH), 162(C-4), 133(C-3\ \&\ 5), 122(C-1)$ and $116(C-2\ \&\ 6)$

122. An organic compoound ($C_9H_{10}O_3$) exhibited the following spectral data:
IR: 34000, 1680 cm^{-1};
^1H NMR: δ 7.8(1H, d, J = 8 Hz), 7.0 (1 H, d, J = 8Hz), 6.5 (1 H, s), 5.8 (1 H, s, D_2O exchangeable), 3.9(3H, s), 2.3 (3 H, s).
The compound is

(a) 

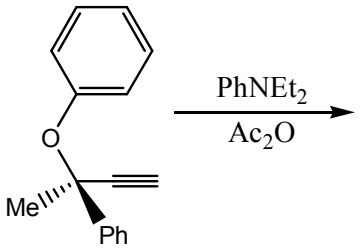
(b) 

(c) 

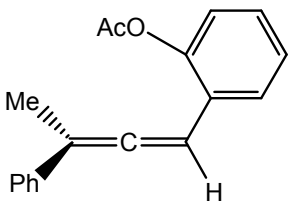
(d) 

123. The $[\alpha]_D$ of a 90% optically pure 2-arylpropanoic acid solution is +135°. On treatment with a base at RT for one hour, $[\alpha]_D$ changed to +120°. The optical purity is reduced to 40% after 3 hours. If so, the optical purity of the solution after 1 hour, and its $[\alpha]_D$ after 3 hours, respectively, would be
(a) 80% and 60° (b) 70% and 40° (c) 80% and 90° (d) 70% and 60°

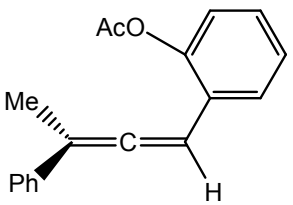
124. In the following pericyclic reaction, the structue of the allene formed and its configuration are



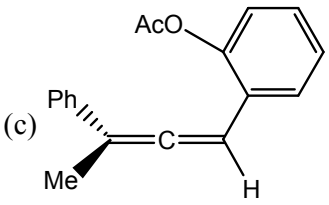
(optically pure)

(a) 

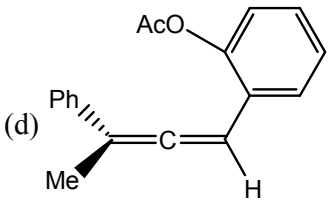
R

(b) 

S

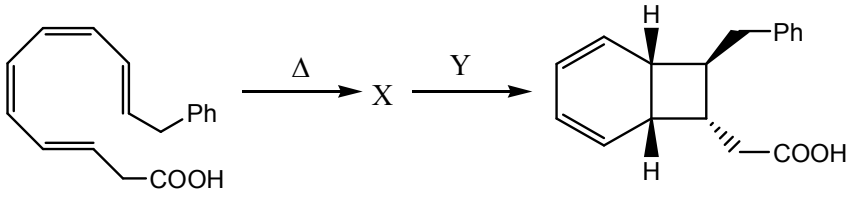
(c) 

R

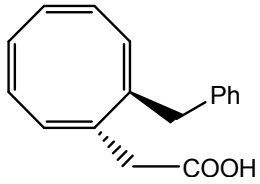
(d) 

S

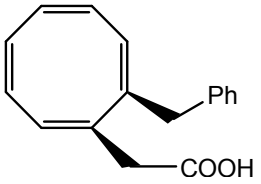
125. In the following sequence of pericyclic reactions X and Y are



X **Y** **X** **Y**

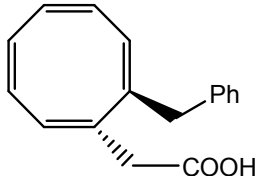
(a) 

hv/DIS

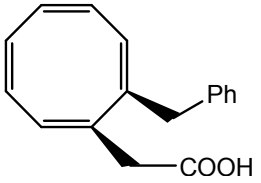
(b) 

hv/CON

X **Y** **X** **Y**

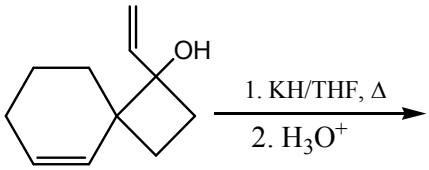
(c) 

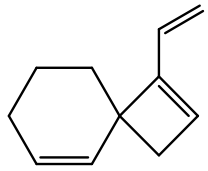
Δ/DIS

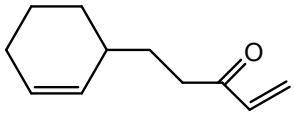
(d) 

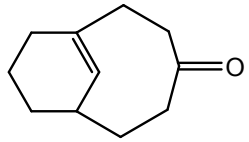
Δ/CON

126. The major product formed in the following reaction is



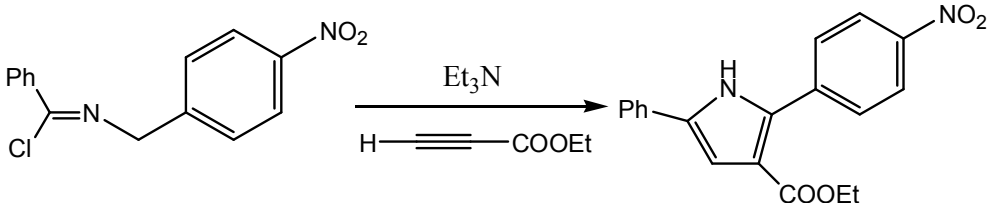
(a) 

(b) 

(c) 

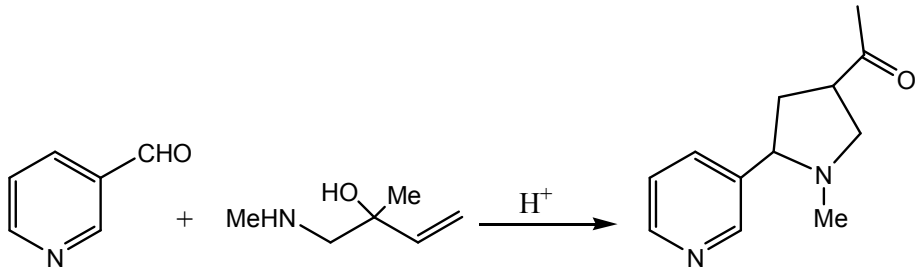
(d) 

127. The following conversion involves



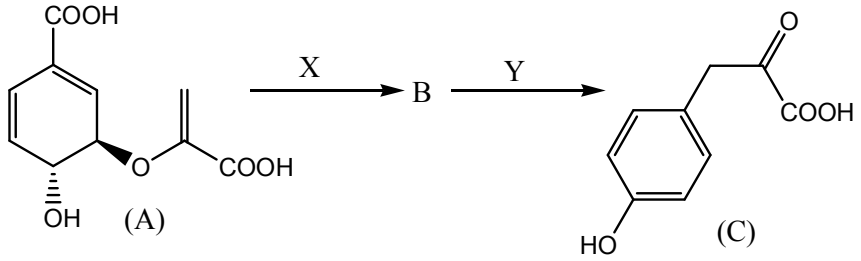
(a) a 1, 3-dipolar species as reactive intermediate, and a cycloaddition.
(b) a carbenium ion as reactive intermediate, and a cycloaddition.
(c) a 1, 3-dipolar species as reactive intermediate, and an aza Witting reaction.
(d) a carbanion as reactive intermediate, and an aza Cope rearrangement.

128. The following transformation involves



(a) an iminium ion, [3, 3]-sigmatropic shift and Mannich reaction.
(b) a nitrenium ion, [3, 3]-sigmatropic shift and Michael reaction.
(c) an iminium ion, [1, 3]-sigmatropic shift and Mannich reaction.
(d) a nitrenium ion, [1, 3]-sigmatropic shift and Michael reaction.

129. With respect to the following biogenetic conversion of chorismic acid (A) to 4-hydroxyphenylpyruvic acid (C), the correct statement is



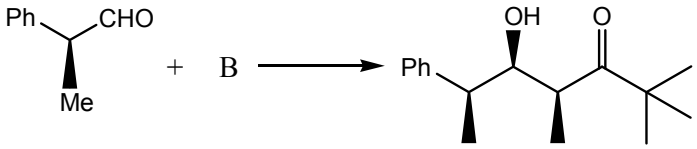
(a) X is Claisen rearrangement; Y is oxidative decarboxylation.
(b) X is Fries rearrangement; Y is oxidative decarboxylation.
(c) X is Fries rearrangement; Y is dehydration.
(d) X is Claisen rearrangement; Y is dehydration.

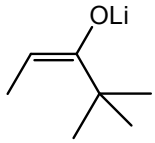
130. Match the following

(i) β-amyrin	(A) alkaloid; secondary alcohol
(ii) squalene	(B) alkaloid, phenol
(iii) morphine	(C) triterpene, secondary alcohol
(iv) ephedrine	(D) acyclic triterpene, polyene

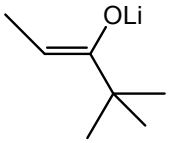
(a) (i)-(C), (ii)-(D), (iii)-(B), (iv)-(A) (b) (i)-(B), (ii)-(A), (iii)-(C), (iv)-(D)
(c) (i)-(C), (ii)-(B), (iii)-(D), (iv)-(A) (d) (i)-(A), (ii)-(D), (iii)-(B), (iv)-(C)

131. In the following reaction, the structure of B, and the mode of addition are

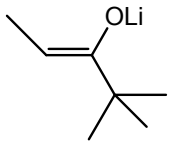


(a) 

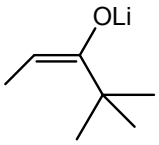
Re-Si facial

(b) 

Re-Re facial

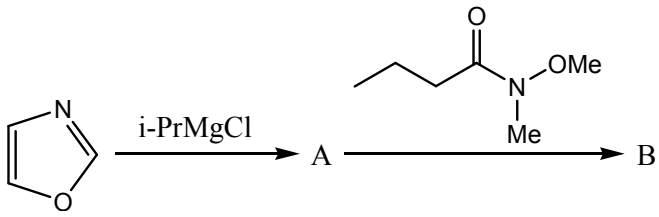
(c) 

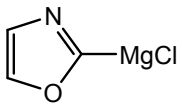
Re-Si facial

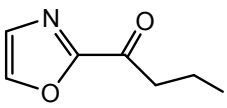
(d) 

Si-Si facial

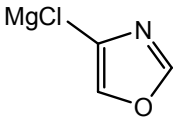
132. In the following reaction A and B are

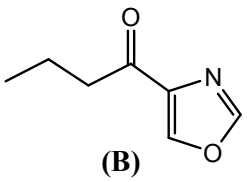


(a) 

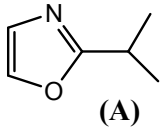
(A)


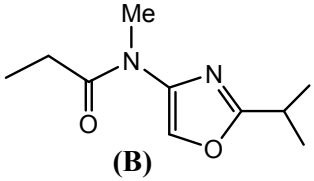
(B)

(b) 

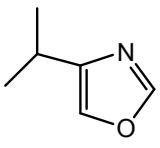
(A)


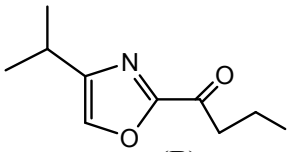
(B)

(c) 

(A)


(B)

(d) 

(A)


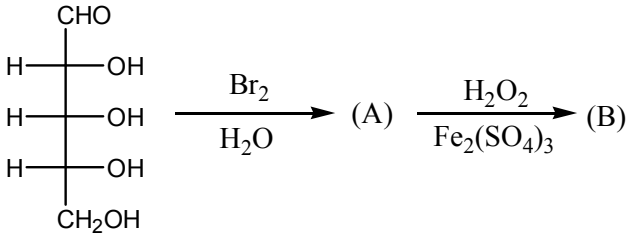
(B)

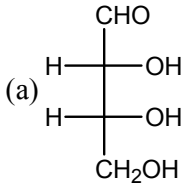
133. Match the following biochemical transformations with coenzymes involved

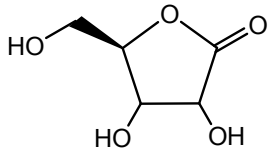
(i) α-ketoglutarate to glutamic acid	(A) tetrahydrofolate
(ii) uridine to thymidine	(B) NADH
(iii) pyruvic acid to acetyl coenzyme A	(C) thiamine pyrophosphate.
	(D) pyridozamine

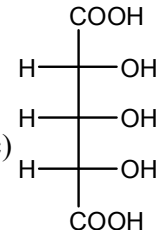
(a) (i)-(D), (ii)-(A), (iii)-(C) (b) (i)-(A), (ii)-(B), (iii)-(D)
(c) (i)-(B), (ii)-(A), (iii)-(C) (d) (i)-(D), (ii)-(B), (iii)-(C)

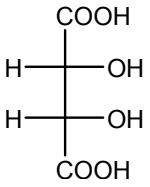
134. The structure of major product B formed in the following reaction sequence is



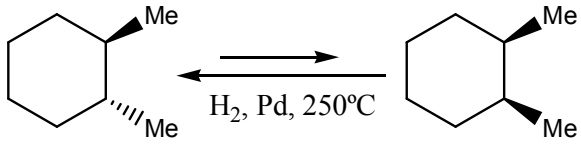
(a) 

(b) 

(c) 

(d) 

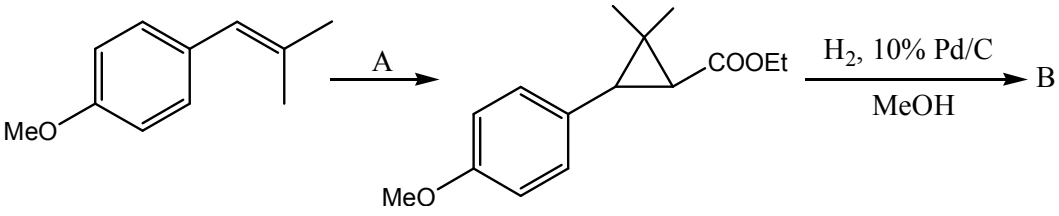
135. Given the energy of each gauche butane interaction is 0.9 kcal/mol, ΔG value of the following reaction is



(e, e conformer)

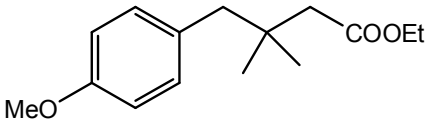
(a) 0.9 kcal/mol (b) 1.8 kcal/mol (c) 2.7 kcal/mol (d) 3.6 kcal/mol

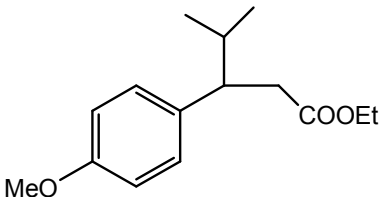
136. In the following reaction, the reagent A and the major product B are



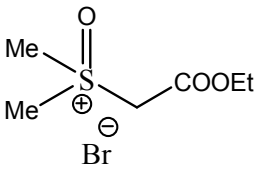
(A) **(B)**

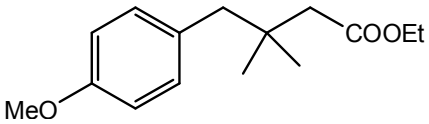
(a) $N_2CHCOOEt$, $Cu(acac)_2$

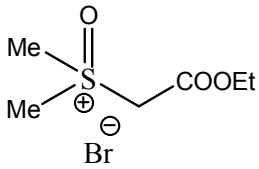


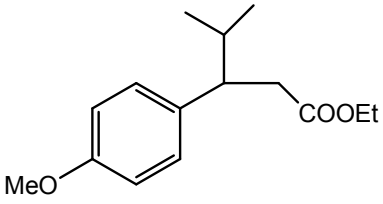


(b) $N_2CHCOOEt$, $Cu(acac)_2$

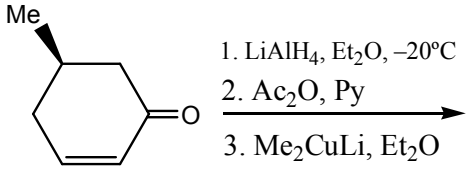
(c) NaH, 

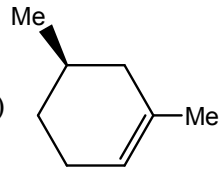


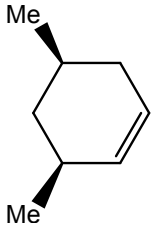
(d) NaH, 

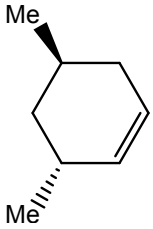


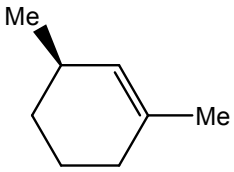
137. The major product formed in the following reaction sequence is



(a) 

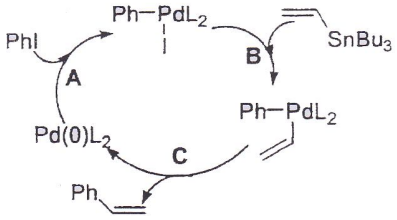
(b) 

(c) 

(d) 

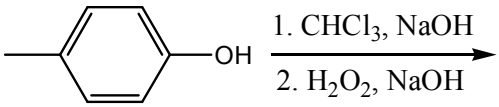
138. 12.0 g of acetophenone on reaction with 76.2 g of iodine in the presence of aq. NaOH gave solid A in 75% yield. Approximate amount of A obtained in the reaction and its structure are
(a) 80 g, Cl_4 (b) 40 g, Cl_4 (c) 60 g, CHI_3 (d) 30 g, CHI_3.

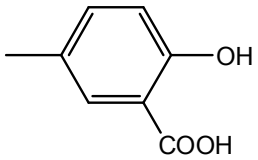
139. Consider the following reaction mechanism

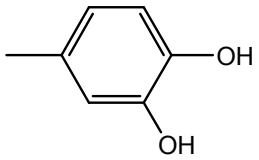


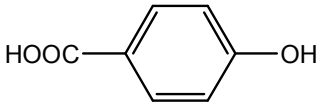
The steps A, B and C, respectively, are
(a) Oxidative addition; transmetallation; reductive elimination.
(b) Oxidative addition; carbopalladation; β-hydride elimination.
(c) Carbopalladation; transmetallation; reductive elimination.
(d) Metal halogen exchange; transmetallation; metal extrusion.

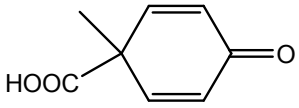
140. The major product formed in the following reaction sequence is



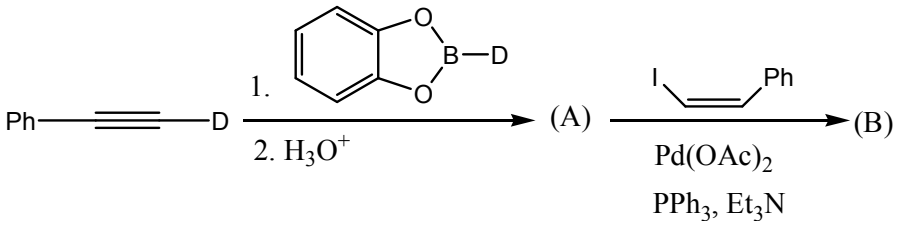
(a) 

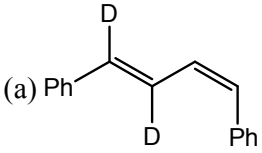
(b) 

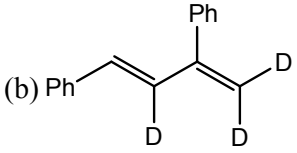
(c) 

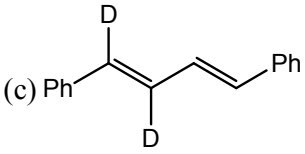
(d) 

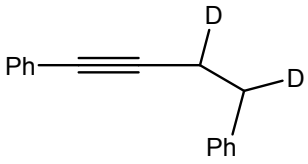
141. The major product B formed in the following reaction sequence is



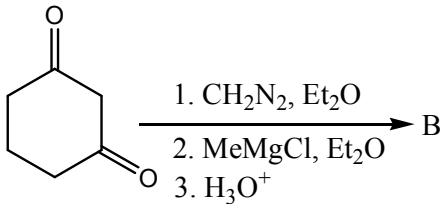
(a) 

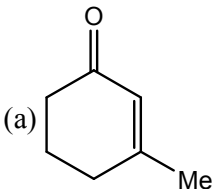
(b) 

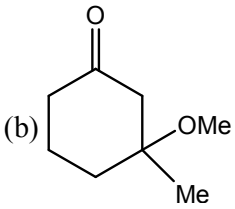
(c) 

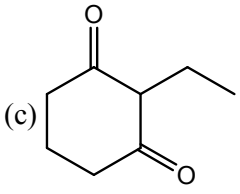
(d) 

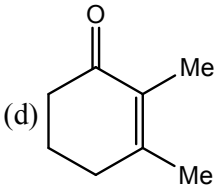
142. The major product B formed in the following reaction sequence is



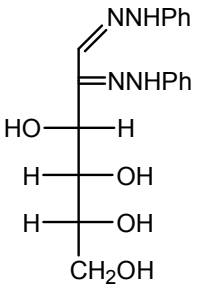
(a) 

(b) 

(c) 

(d) 

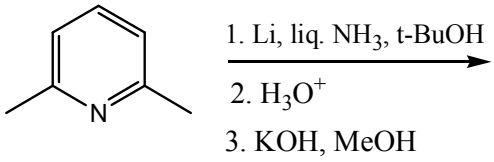
143. The osazone A could be obtained from

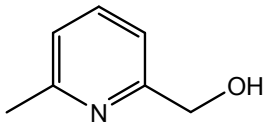


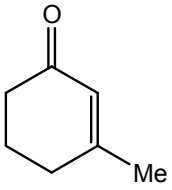
(A)

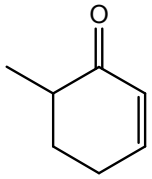
(a) glucose and mannose
(b) mannose and galactose
(c) gulcose and fructose
(d) galactose and fructose

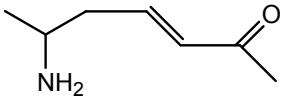
144. The major product formed in the following reaction is:



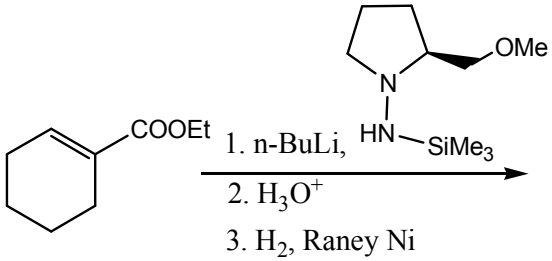
(a) 

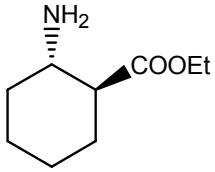
(b) 

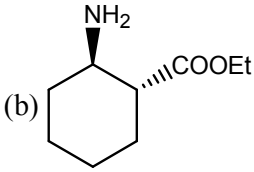
(c) 

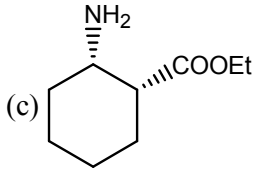
(d) 

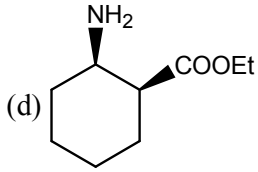
145. In the following enantioselective reaction, the major product formed is



(a) 

(b) 

(c) 

(d) 

NET-JRF DEC-2012
Chemical Sciences (Answer Key)

Part-A:

1. (a)	**2. (b)**	**3. (d)**	**4. (c)**	**5. (c)**	**6. (d)**	**7. (c)**
8. (d)	**9. (a)**	**10. (a)**	**11. (b)**	**12. (a)**	**13. (b)**	**14. (d)**
15. (c)	**16. (c)**	**17. (c)**	**18. (c)**	**19. (a)**	**20. (a)**	

Part-B:

21. (a)	**22. (d)**	**23. (b)**	**24. (d)**	**25. (d)**	**26. (c)**	**27. (a)**
28. (d)	**29. (b)**	**30. (a)**	**31. (b)**	**32. (d)**	**33. (c)**	**34. (a)**
35. (a)	**36. (d)**	**37. (a)**	**38. (d)**	**39. (c)**	**40. (c)**	**41. (b)**
42. (c)	**43. (d)**	**44. (d)**	**45. (a)**	**46. (b)**	**47. (a)**	**48. (b)**
49. (c)	**50. (a)**	**51. (d)**	**52. (b)**	**53. (b)**	**54. (a)**	**55. (d)**
56. (b)	**57. (c)**	**58. (d)**	**59. (b)**	**60. (b)**	**61. (a)**	**62. (a)**
63. (d)	**64. (d)**	**65. (a)**	**66. (b)**	**67. (d)**	**68. (b)**	**69. (c)**
70. (d)						

Part-C:

71. (d)	**72. (a)**	**73. (c)**	**74. (c)**	**75. (b)**	**76. (b)**	**77. (d)**
78. (a)	**79. (a)**	**80. (a)**	**81. (d)**	**82. (a)**	**83. (a)**	**84. (b)**
85. (a)	**86. (d)**	**87. (a)**	**88. (a)**	**89. (d)**	**90. (b)**	**91. (d)**
92. (a)	**93. (b)**	**94. (b)**	**95. (a)**	**96. (c)**	**97. (c)**	**98. (a)**
99. (c)	**100. (c)**	**101. (b)**	**102. (b)**	**103. (a)**	**104. (c)**	**105. (c)**
106. (c)	**107. (a)**	**108. (b)**	**109. (c)**	**110. (b)**	**111. (c)**	**112. (d)**
113. (d)	**114. (c)**	**115. (b)**	**116. (b)**	**117. (a)**	**118. (c)**	**119. (a)**
120. (d)	**121. (b)**	**122. (d)**	**123. (a)**	**124. (a)**	**125. (c)**	**126. (c)**
127. (a)	**128. (a)**	**129. (a)**	**130. (a)**	**131. (c)**	**132. (a)**	**133. (a)**
134. (a)	**135. (b)**	**136. (a)**	**137. (c)**	**138. (d)**	**139. (a)**	**140. (b)**
141. (a)	**142. (a)**	**143. (a)**	**144. (b)**	**145. (a)**		